

Ox. #	Anhydride	Conjugate Acid Formula	Conjugate Acid Name	Acid/Base Formula	Acid/Base Name	Conjugate Base Formula	Conjugate Base Name
<b>Boron</b>							
3	B <sub>2</sub> O <sub>3</sub>			H <sub>3</sub> BO <sub>3</sub>	Boric acid	BO <sub>3</sub> <sup>3-</sup>	Orthoborate
3	B <sub>2</sub> O <sub>3</sub>			H <sub>2</sub> B <sub>4</sub> O <sub>7</sub>	Tetraboric acid	B <sub>4</sub> O <sub>7</sub> <sup>2-</sup>	Tetraborate
				HF <sub>4</sub>	Fluoboric acid	BF <sub>4</sub> <sup>-</sup>	Fluoborite
<b>Carbon</b>							
-1				HCN	Hydrocyanic acid	CN <sup>-</sup>	Cyanide
				CH <sub>2</sub> O <sub>2</sub>	Formic acid, Methanoic acid	CHO <sub>2</sub> <sup>-</sup>	formate, methanoate
				CH <sub>3</sub> CO <sub>2</sub> H	Acetic acid, Ethanoic acid	CH <sub>3</sub> CO <sub>2</sub> <sup>-</sup>	acetate, ethanoate
				C <sub>2</sub> H <sub>5</sub> CO <sub>2</sub> H	Propionic acid, Propanoic acid	C <sub>2</sub> H <sub>5</sub> CO <sub>2</sub> <sup>-</sup>	propionate, propanoate
3				H <sub>2</sub> C <sub>2</sub> O <sub>4</sub>	Oxalic acid	C <sub>2</sub> O <sub>4</sub> <sup>2-</sup>	Oxalate
4	CO <sub>2</sub>			H <sub>2</sub> CO <sub>3</sub>	Carbonic acid	CO <sub>3</sub> <sup>2-</sup>	Carbonate
6				H <sub>2</sub> C <sub>2</sub> O <sub>4</sub>	Hydroperoxyformic acid	CO <sub>4</sub> <sup>2-</sup>	Peroxy carbonate
<b>Nitrogen</b>							
-3		NH <sub>4</sub> <sup>+</sup>	Ammonium ion	H <sub>3</sub> N, NH <sub>3</sub>	Ammonia (weak base)	N <sup>3-</sup>	Nitride
1	N <sub>2</sub> O			H <sub>2</sub> N <sub>2</sub> O <sub>2</sub>	Hyponitrous acid	N <sub>2</sub> O <sub>2</sub> <sup>2-</sup>	Hyponitrite
3	N <sub>2</sub> O <sub>3</sub>			HNO <sub>2</sub>	Nitrous acid	NO <sub>2</sub> <sup>-</sup>	Nitrite
5	N <sub>2</sub> O <sub>5</sub>			HNO <sub>3</sub>	Nitric acid	NO <sub>3</sub> <sup>-</sup>	Nitrate
7	N <sub>2</sub> O <sub>7</sub>			HNO <sub>4</sub>	Peroxynitric acid	NO <sub>4</sub> <sup>-</sup>	Peroxynitrate
<b>Oxygen</b>							
-1				H <sub>2</sub> O <sub>2</sub>	Hydrogen peroxide (very basic)	O <sub>2</sub> <sup>2-</sup>	Peroxide
-2				H <sub>3</sub> O <sup>+</sup>	Hydronium ion	H <sub>2</sub> O	Water
-2		H <sub>3</sub> O <sup>+</sup>	Hydronium ion	H <sub>2</sub> O	Water	OH <sup>-</sup>	Hydroxide
-2		H <sub>2</sub> O	Water	OH <sup>-</sup>	Hydroxide	O <sup>2-</sup>	Oxide

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<b>Fluorine</b>							
-1				HF	Hydrofluoric acid	F <sup>-</sup>	Fluoride
<b>Phosphorus</b>							
-3				PH <sub>3</sub>	Phosphine (low water solubility)	P <sup>3-</sup>	Phosphide
1	P <sub>2</sub> O			H <sub>3</sub> PO <sub>2</sub>	Hypophosphorous acid	PO <sub>2</sub> <sup>3-</sup>	Hypophosphite
3	P <sub>2</sub> O <sub>3</sub>			H <sub>3</sub> PO <sub>3</sub>	Phosphorous acid	PO <sub>3</sub> <sup>3-</sup>	Phosphite
5	P <sub>2</sub> O <sub>5</sub> (P <sub>4</sub> O <sub>10</sub> )			H <sub>3</sub> PO <sub>4</sub>	Phosphoric acid	PO <sub>4</sub> <sup>3-</sup>	Phosphate
7				H <sub>3</sub> PO <sub>5</sub>	Peroxymonophosphoric acid	PO <sub>5</sub> <sup>3-</sup>	Peroxymonophosphate
<b>Sulfur</b>							
-2				H <sub>2</sub> S	Hydrosulfuric acid	S <sup>2-</sup>	Sulfide
2	SO, S <sub>2</sub> O <sub>2</sub>	H <sub>2</sub> SO <sub>2</sub>		HSO <sub>2</sub> <sup>-</sup>	Hyposulfuric acid	SO <sub>2</sub> <sup>2-</sup>	hyposulfite, sulfoxylate
2	SO, S <sub>2</sub> O <sub>2</sub>			H <sub>2</sub> S <sub>2</sub> O <sub>3</sub>	Thiosulfuric acid	S <sub>2</sub> O <sub>3</sub> <sup>2-</sup>	thiosulfate
4	SO <sub>2</sub>			H <sub>2</sub> SO <sub>3</sub>	Sulfurous acid	SO <sub>3</sub> <sup>2-</sup>	sulfite
				CH <sub>3</sub> SO <sub>2</sub> O H	Methanesulfonic acid	CH <sub>3</sub> SO <sub>3</sub> <sup>-</sup>	methanesulfonate
6	SO <sub>3</sub>			H <sub>2</sub> SO <sub>4</sub>	Sulfuric acid	SO <sub>4</sub> <sup>2-</sup>	sulfate
6	S <sub>2</sub> O <sub>5</sub> Cl <sub>2</sub>			HSO <sub>3</sub> Cl	Chlorosulfuric acid (dec. in H <sub>2</sub> O)	SO <sub>3</sub> Cl <sup>-</sup>	chlorosulfate
8				H <sub>2</sub> SO <sub>5</sub>	Peroxymonosulfuric acid	SO <sub>5</sub> <sup>2-</sup>	peroxymonosulfate
<b>Chlorine</b>							
-1				HCl	Hydrochloric acid	Cl <sup>-</sup>	chloride
1	Cl <sub>2</sub> O			HClO	Hypochlorous acid	ClO <sup>-</sup>	hypochlorite
3	Cl <sub>2</sub> O <sub>3</sub> (hyp.)			HClO <sub>2</sub>	Chlorous acid	ClO <sub>2</sub> <sup>-</sup>	chlorite
5	Cl <sub>2</sub> O <sub>5</sub> (hyp.)			HClO <sub>3</sub>	Chloric acid	ClO <sub>3</sub> <sup>-</sup>	chlorate
7	Cl <sub>2</sub> O <sub>7</sub>			HClO <sub>4</sub>	Perchloric acid	ClO <sub>4</sub> <sup>-</sup>	perchlorate

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<b>Chromium</b>							
6	CrO <sub>3</sub>			H <sub>2</sub> CrO <sub>4</sub>	Chromic acid	CrO <sub>4</sub> <sup>2-</sup>	chromate
6	CrO <sub>3</sub>			H <sub>2</sub> CrO <sub>4</sub>	Chromic acid	Cr <sub>2</sub> O <sub>7</sub> <sup>2-</sup>	dichromate
<b>Manganese</b>							
6	MnO <sub>3</sub>			NA	Disproportionates w/o strong base	MnO <sub>4</sub> <sup>2-</sup>	manganate
7	Mn <sub>2</sub> O <sub>7</sub>			HMnO <sub>4</sub>	Permanganic acid	MnO <sub>4</sub> <sup>-</sup>	permanganate
<b>Arsenic</b>							
-3				AsH <sub>3</sub>	Arsenic trihydride, arsine	As <sup>3-</sup>	arsenide
3	As <sub>2</sub> O <sub>3</sub>			H <sub>3</sub> AsO <sub>3</sub>	Arsenous acid	AsO <sub>3</sub> <sup>3-</sup>	arsenite
5	As <sub>2</sub> O <sub>5</sub>			H <sub>3</sub> AsO <sub>4</sub>	Arsenic acid	AsO <sub>4</sub> <sup>3-</sup>	arsenate
<b>Selenium</b>							
-2				H <sub>2</sub> Se	Hydroselenic acid	Se <sup>2-</sup>	selenide
4	SeO <sub>2</sub>			H <sub>2</sub> SeO <sub>3</sub>	Selenous acid	SeO <sub>3</sub> <sup>2-</sup>	selenite
6	SeO <sub>3</sub>			H <sub>2</sub> SeO <sub>4</sub>	Selenic acid	SeO <sub>4</sub> <sup>2-</sup>	selenate
<b>Bromine</b>							
-1				HBr	Hydrobromic acid	Br <sup>-</sup>	bromide
1	Br <sub>2</sub> O			HBrO	Hypobromous acid	BrO <sup>-</sup>	hypobromite
3	Br <sub>2</sub> O <sub>3</sub>			HBrO <sub>2</sub>	Bromous acid	BrO <sub>2</sub> <sup>-</sup>	bromite
5	Br <sub>2</sub> O <sub>5</sub>			HBrO <sub>3</sub>	Bromic acid	BrO <sub>3</sub> <sup>-</sup>	bromate
7				HBrO <sub>4</sub>	Perbromic acid (unstable)	BrO <sub>4</sub> <sup>-</sup>	perbromate
<b>Tellurium</b>							
-2				H <sub>2</sub> Te	Hydrotelluric acid	Te <sup>2-</sup>	telluride
4	TeO <sub>2</sub>			H <sub>2</sub> TeO <sub>3</sub>	Tellurous acid	TeO <sub>3</sub> <sup>2-</sup>	tellurite
6	TeO <sub>3</sub>			H <sub>2</sub> TeO <sub>4</sub>	Telluric acid	TeO <sub>4</sub> <sup>2-</sup>	tellurate

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<b>Iodine</b>							
-1				HI	Hydroiodic acid	I <sup>-</sup>	iodide
1	I <sub>2</sub> O			HIO	Hypoiodous acid	IO <sup>-</sup>	hypoiodite
3				HIO <sub>2</sub>	Iodous acid	IO <sub>2</sub> <sup>-</sup>	iodite
5	I <sub>2</sub> O <sub>5</sub>			HIO <sub>3</sub>	Iodic acid	IO <sub>3</sub> <sup>-</sup>	iodate
7				HIO <sub>4</sub>	Periodic acid (unstable)	IO <sub>4</sub> <sup>-</sup>	periodate

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